Ionic Bonding Poster

| Component | 4 | $\mathbf{3}$ | $\mathbf{2}$ | $\mathbf{1}$ | 0 |
| :--- | :--- | :--- | :--- | :--- | :--- |
| Layout | Poster is clearly titled <br> with the correct <br> formula of the ionic <br> compound and <br> compound name based <br> on the two elements <br> assigned to the group | Correct formula of the <br> ionic compound and <br> compound name can <br> be found somewhere <br> on the poster | The formula of the <br> ionic compound and <br> compound name are <br> present based on the <br> two elements assigned <br> to the group. There is <br> an error in one or <br> both. | Poster is missing <br> either the compound <br> name or formula. | Not Present |
| Definitions | The terms ionic bond, <br> cation, anion, and <br> crystal lattice are <br> defined clearly on the <br> poster | 1 of the terms is <br> missing | 2 of the terms are <br> missing | 3 of the terms are <br> missing | Not Present |
| Shell Models | An accurate shell <br> model is drawn for <br> both assigned <br> elements. The transfer <br> of electrons between <br> the atom(s) of each <br> element are shown. | Accurate shell models <br> are drawn, but the <br> transfer of electrons is <br> incorrect | Errors in the shell <br> models and/or <br> electron transfer not <br> shown | Attempt | Not Present |
| Change in atomic <br> radius | The initial drawing of <br> each element is to <br> scale based on its <br> relative position on the <br> periodic table. Poster <br> clearly shows how the <br> size of each atom <br> changes when it <br> becomes an ion | Poster clearly shows <br> how the size of each <br> atom changes when it <br> becomes an ion | Poster incorrectly <br> shows how the size of <br> each atom changes <br> when it becomes an <br> ion | Only 1 atom is shown | Not Present |
| Crystal Lattice | Poster correctly shows <br> how the ions of each <br> element would be <br> arranged in a solid. <br> The correct ratio <br> between cations and <br> anions is shown. | Poster correctly shows <br> how the ions of each <br> element would be <br> arranged in a solid. <br> The incorrect ratio <br> between cations and <br> anions is shown. | Poster incorrectly <br> shows how the ions of <br> each element would be <br> arranged in a solid. | Attempt | Not Present |

$\qquad$

