



# Unit 4: Chemical Reactions

Lesson 11: Can we name it?

**Guiding Question:** Explain how to name binary covalent compounds and how it differs from ionic compounds.

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**Do Now:**

- Draw a structural formula for SF<sub>2</sub>. Draw and name the molecular geometry.

# Notes

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- Binary molecular covalent compounds contain 2 elements that are both nonmetals.
- Nonmetals can combine in a variety of ways so we need to indicate how many of each nonmetal are present in a compound using prefixes (see the table on pg. 25 for prefixes)

# Notes

- When naming use the following rules:
  - Prefix-first element + prefix-second element-ide ending
  - Never use the prefix mono on the first element

Ex:



Sulfur Difluoride

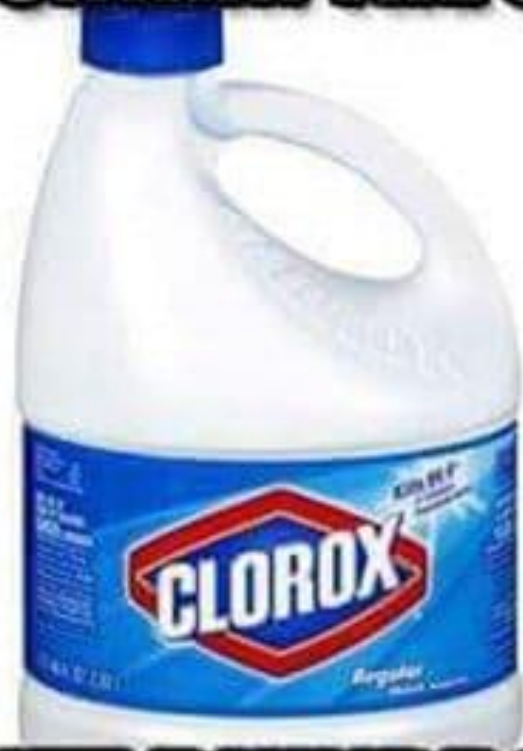


Diphosphorus Pentasulfide



Diboron Monosilicide

**THESE PRODUCTS SHOULD NOT  
CONTAIN THE SAME CHEMICALS**



**YET DIHYDROGEN MONOXIDE  
CAN BE FOUND IN BOTH**

# Closure

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- Answer Guiding Question on page 15
- Quiz #3 next block (3/7 and 3/8)
  - Quiz #3 review due Friday, 3/9 (you will get this tomorrow)
- Achieve 3000: The Car that Runs on Chocolate due Friday, 3/9 at 11:59pm