



# Unit 4: Chemical Reactions

Lesson 14: Think (Electro)negatively

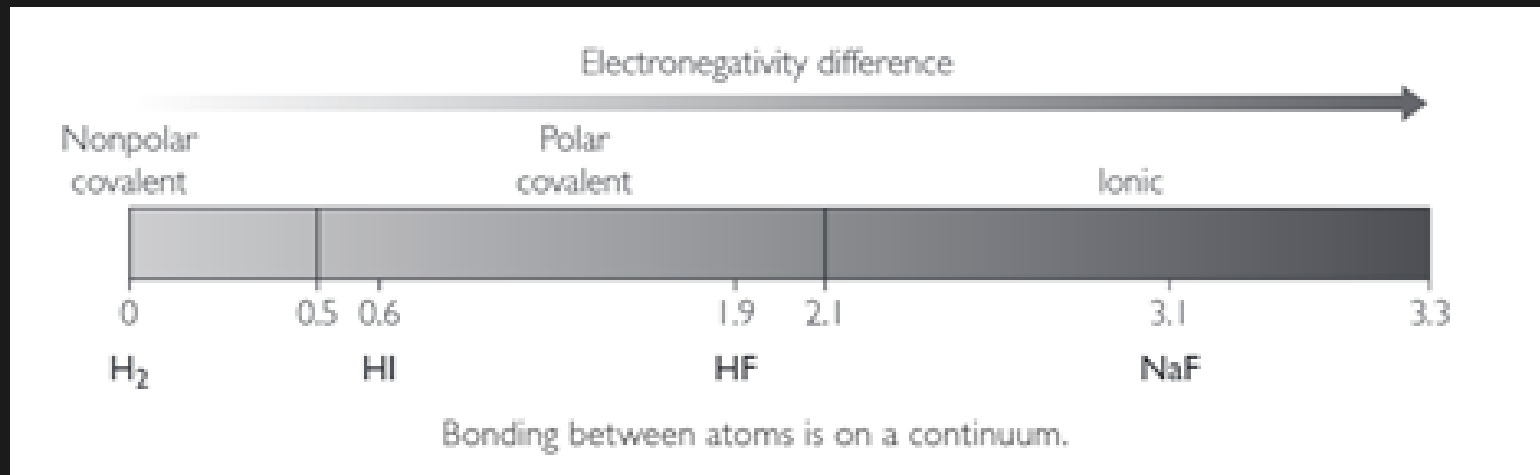
**Guiding Question:** How can we predict the type of bond in a molecule using electronegativity values?

**Do Now:**

- What type of bonding do you think is present in the following molecules? Explain your reasoning.
  1.  $\text{CH}_4$
  2.  $\text{H}_2\text{O}$
  3.  $\text{K}_2\text{S}$

# Notes

Numerical differences in electronegativity can help us predict the type of bond that will form between two atoms.



# Check in

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Is the bond between potassium chloride, nonpolar, polar or ionic? Explain.

What happens to the valence electrons in KCl?

# Closure

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- Answer Guiding Question on page 13.
- No homework this week (yay!)
  - Until our Performance Task
- Workbook 4.3 due at the block (3/14 & 3/15)
- Achieve 3000: “The Missouri Gets a Makeover” due Friday, 3/23 at 11:59pm.