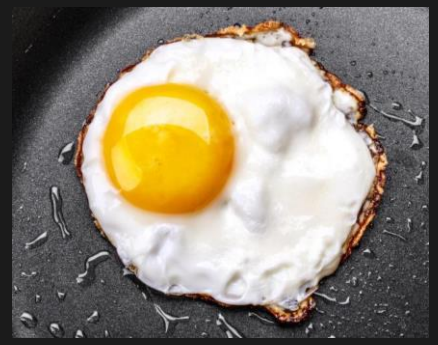
Unit 4: Chemical Reactions

Lesson 15: Spare Change

Guiding Question: Explain the differences between physical and chemical changes.

Do Now:

How could you group these different phenomena?



Notes

- Chemical equations represent changes in matter using symbols and formulas. They also indicate what state of matter each chemical is in. Recall (s) = solid, (l) = liquid , (g) = gas, and (aq) =aqueous (dissolved in water).
- These equations allow you to see what is happening to matter at an atomic level.
- They do not tell you how fast the change will happen, if it will happen at all, nor all of the things you may observe (color changes, heat, fire, etc.)

Notes

- Physical Change: A change in matter in which a substance changes form but not identity.
- Examples:
  - Torn Paper
  - Chopped Wood
  - Salt in Water

Notes

- Chemical Change: A change in matter that results in the formation of a new substance or substances with new properties.
- Examples:
  - Fried Egg
  - Moldy Bread

Check in

- What would you expect to observe in the following reaction based on the information given?

$$CoCl_2(aq) + Ca(OH)_2(aq) \rightarrow Co(OH)_2(s) + CaCl_2(aq)$$

Closure

- Achieve 3000: “The Missouri gets a Makeover” due Friday 3/23 at 11:59pm
- Turn in you Chromatography PT to the bin on your way out
- Benchmark #3 on 3/28 & 3/29
- Homework #5 due Friday, 3/23
  - This will be handed out tomorrow (Tuesday)